

II ()

11

1. 264 $\text{Hg}_2(\text{NO}_3)_2$ 20 %
 (I) ? 6 %
 1. ?
 2. ?

- 20

2. 150 44,8
 (. .).
 2,16 .
 1.
 2.

- 20

3. :
 $\xrightarrow[\text{t}^0]{\text{Cl}_2, \text{h}}$ X_2 C_2H_6 X_3 $\xrightarrow[\text{Cl}_2]{\text{Na}}$ C_3H_8 X_4 $\xrightarrow[\text{2NH}_3]{\text{Br}_2, \text{t}^0}$ X_5 $\xrightarrow[\text{H}_2\text{SO}_4]{\text{KMnO}_4, \text{Cl}_2}$ X_6 X_7 X_8
 1 9 3 10 11 12

- 20

4. (V)
 $\text{P}_4() + 6\text{Cl}_2() = 4 \text{P}_3()$
 $\text{PCl}_3() + \text{Cl}_2() = \text{PCl}_5()$
 $^0 = -1224 /$ (I)
 $^0 = -93 /$ (II)

- 20

5. :
 1. , ?
 2. .
 3. .

- 20