

10

:

10-1 (16)

600 CuSO₄
 CuO: $n(\text{CuO}) = n(\text{CuSO}_4) = 1,5/80 = 0,019$
 - CuO.

260 CuSO₄·xH₂O CuSO₄: $n(\text{CuSO}_4) = n(\text{CuSO}_4 \cdot x\text{H}_2\text{O})$
 $= 0,019$. $M(\text{CuSO}_4 \cdot x\text{H}_2\text{O}) = m/n = 4,01/0,019 = 211$ / ; $(160 + 18x) = 211$; $x = 3$.
 - CuSO₄.

95 CuSO₄·yH₂O CuSO₄·3H₂O: $n(\text{CuSO}_4 \cdot y\text{H}_2\text{O}) = n(\text{CuSO}_4 \cdot 3\text{H}_2\text{O})$
 $= 0,019$. $M(\text{CuSO}_4 \cdot y\text{H}_2\text{O}) = m/n = 4,69/0,019 = 247$ / ; $(160 + 18) = 247$; $y = 5$.
 - CuSO₄·3H₂O. -

CuSO₄·5H₂O.

- 8 . : - 2 .

10 - 2 (9)

$$2^2 + 2^2 \rightarrow 2^2$$

$$4 + 2^2 \rightarrow 2^2 + 2^2$$

$$V(\text{H}_2) = x, \quad V(\text{CH}_4) = y.$$

$$2y \quad , \quad x + y = 0,5x + 2y$$

$$x = 2y.$$

0,5 ,

1 : 2.

: - 9 .

10 - 3 (20)

:

$$\frac{m}{n} -$$

(, ,),

$$x -$$

b u

$$207 / \quad 64 / ,$$

$$207n -$$

$$64n -$$

$$xn -$$

$$(207n - xn) -$$

$$(xn - 64n) -$$

$$\frac{n(207 - x)}{m} = 0,19 -$$

$$\frac{n(x - 64)}{m} = 0,096 -$$

$$\begin{cases} \frac{n(207 - x)}{m} = 0,19 \\ \frac{n(x - 64)}{m} = 0,096 \end{cases}$$

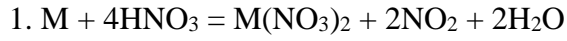
:

$$\frac{207 - x}{x - 64} = \frac{0,19}{0,096}$$

$$x = 112,$$

: -20 .

10 - 4 (20)

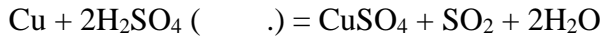
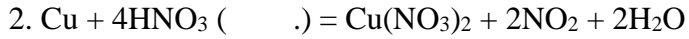


$$m(NO_2) = 10 + 4,38 = 14,38$$

$$10/M(M) = 14,38/92$$

$$M(M) = 64 /$$

- .



: - 2 (6),
- 14 .

10 - 5 (15)

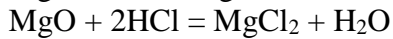
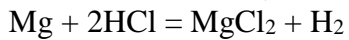
$$500 \cdot 1,10 \cdot x / (500 \cdot 1,10 \cdot x + 2500) = 0,04$$

$$x = 0,189$$

$$w(HCl) = 18,9 \%$$

: - 15 .

10-6 (20)



:

$$3,74 \cdot 740/298 = V \cdot 760/273$$

$$V = 3,34$$

$$m(Mg) = 3,34 \cdot 24/22,4 = 3,58$$

$$w(Mg) = 3,58/5,1 \cdot 100\% = 70,2\%$$

: -16 ,
- 8 ,

2 (4).