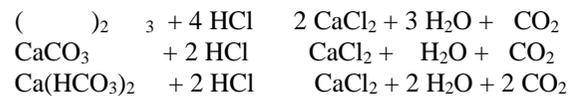
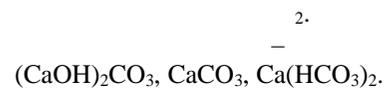


2

9-6 (15 \quad)



1/4 : 1/2 : 1 \quad 1 : 2 : 4

6

-9

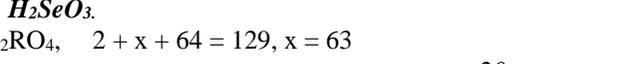
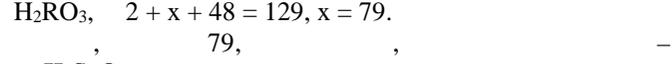
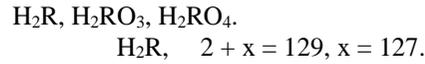
9-1 (20) : 100

1. $n(\text{NaOH}) = 2 \cdot 0,03 = 0,06$.

2. $n(\text{VI}) = 0,06/2 = 0,03$ (, , ,)

3. $(\text{VI}) = 3,87/0,03 = 129 /$.

4. 129 :

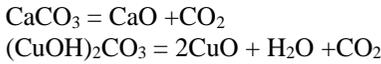


9-2 (20)

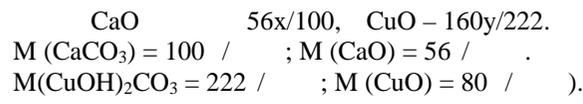


$n(\text{H}_2) = 5,56 / 22,4 = 0,25$
 $n(\text{A}) / 2 = n(\text{H}_2) = 0,5$
 $n(\text{A}) = m(\text{A}) / M(\text{A}) = 10 / M(\text{A})$
 $10 / (M(\text{A}) \cdot x) = 0,5$
 $M(\text{A}) \cdot x = 10 / 0,5 = 20$

9-3 (15)

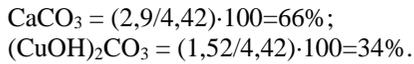


$4,42$. $-x,$ $-y,$ $x+y=$



$\begin{cases} x + y = 4,42 \\ (56x/100) + (160y/222) = 2,72 \end{cases}$
 $x=2,9; \quad y=1,52, \dots$ $2,9 ,$

$1,52$.



9-4 (10)

- 1) $\text{N}_2 + 3\text{H}_2 = 2\text{NH}_3$
- 2) $2\text{NH}_3 + \text{H}_2\text{SO}_4 = (\text{NH}_4)_2\text{SO}_4$
- 3) $(\text{NH}_4)_2\text{SO}_4 + \text{BaCl}_2 = \text{BaSO}_4 + 2\text{NH}_4\text{Cl}$
- 4) $2\text{NH}_4\text{Cl} + \text{Ca}(\text{OH})_2 = \text{CaCl}_2 + 2\text{NH}_3 + 2\text{H}_2\text{O}$

9-5 (20)

