

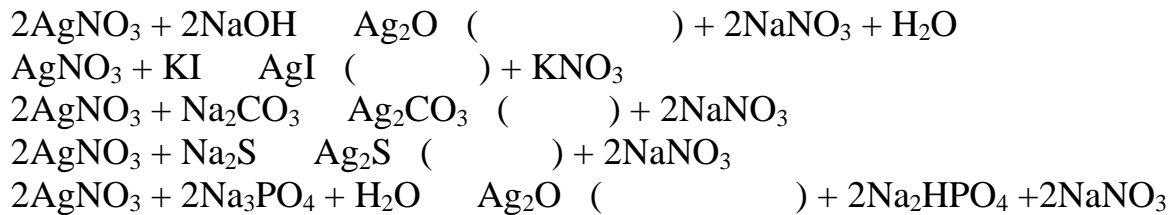
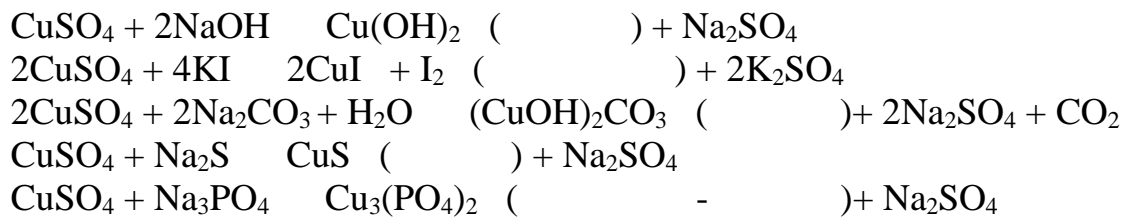
. . .

1.

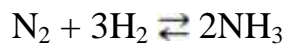
1.1.

	NaOH	KI	Na ₂ CO ₃	Na ₂ S	Na ₃ PO ₄
CuSO ₄					-
AgNO ₃					

1.2.



2.



. , - .
x . ,

	N ₂	H ₂	NH ₃
	5	5	0
	-1/3x	-x	+2x/3
	5 - 1/3x	5 - x	2x/3

$5 - 1/3x + 5 - x + 2x/3 = 8$,
 $2/3 \cdot 3 = 2$.
 $x = 3$.

$$\eta = \frac{V_{\text{факт}}}{V_{\text{теор}}} 100\%$$

$$\eta = \frac{2 \text{ л}}{\frac{10}{3} \text{ л}} 100\% = 60\%$$

3.

3.1.



$$n(\text{K}) = \frac{m(\text{K})}{M(\text{K})} = \frac{1.56\text{г}}{39 \frac{\text{г}}{\text{моль}}} = 0.04 \text{ моль}$$

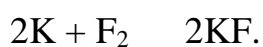
$$n(\text{K}_m\text{X}_n) = \frac{n(\text{K})}{m} = \frac{0.04}{m} \text{ моль}$$

$$M(\text{K}_m\text{X}_n) = \frac{m(\text{K}_m\text{X}_n)}{n(\text{K}_m\text{X}_n)} = \frac{2.32}{\frac{0.04}{m}} = 58m$$

$$M(\text{X}_n) = M(\text{K}_m\text{X}_n) - M(\text{K}_m) = 58m - 39m = 19m$$

$$m = 1 \quad \text{KF} \quad M(\text{X}) = 19 \text{ г/моль}$$

$$m > 13, \quad \text{X} - \text{F}_2, \quad \text{Y} - \text{KF}$$



3.2.

$$V(\text{F}_2) = n(\text{F}_2) \times V_m = \frac{1}{2} n(\text{K}) \times V_m = 0.02 \text{ моль} \times 22.4 \frac{\text{л}}{\text{моль}} = 448 \text{ мл}$$

4.



$$250 \cdot 0.25 = 62.5 \text{ g}$$

$$x - 0.1x, \quad (0.1x)/142$$

$$261 \cdot (0.1x)/142, \quad (0.1x)/142$$

$$62.5 - 261 \cdot (0.1x)/142$$

$$233 \cdot (0.1x)/142$$

$$250 + x - 233 \cdot (0.1x)/142$$

$$\omega(\text{Ba}(\text{NO}_3)_2) = \frac{62.5 - 261 \frac{0.1x}{142}}{250 + x - 233 \frac{0.1x}{142}} = 0.2$$

$$x = 35.6 \text{ g}$$

1	
(5) - 1	5
(10) - 2	20
	25

2	
	2
	3
	10
	5
	5
	25

3	
	4
	3
	5
	5
	3
	5
	25

4	
	3
	2
,	4
	4
	4
	4
	4
	25