

10

1. 10

1	2	3	4	5	6	7	8	9	10
4	4	3	4	4	3	3	2	2	2

2. 9

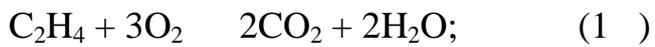
- 1) $2\text{Al} + 6\text{HCl} = \text{AlCl}_3 + 3\text{H}_2$ (1)
 $\frac{54}{67,2}$
- 2) $x = (54 \cdot 6,72) / 67,2 = 5,4$; $m(\text{H}_2) = 5,4$ (2)
- 3) $3\text{Me} + 4\text{HNO}_3 = 3\text{MeNO}_3 + \text{NO} + 2\text{H}_2\text{O}$ (1)
 $\frac{y}{30}$
- $y = (5,4 \cdot 30) / 0,5 = 324$ (1)
- 4) $M(\text{Me}) = 324 / 3 = 108$ / - (1)
- 5) $(\text{Al}) = 5,4 / 27 = 0,2$ (1)
 $(\text{Ag}) = 5,4 / 108 = 0,05$ (1)
- 6) $(\text{Al}) : (\text{Ag}) = 0,2 : 0,05 = 4:1$ (1)

3. 9

- 1) $44,8$, $22,4$. (1)
 $1013 / 101,3 = 10$. (1)
- 2) 10 . (1)
- 3) $202,6 / 101,3 = 2$. (1)
- 4) 2 . (1)
- 5)
- $\text{NH}_3(\text{g}) + \text{HCl}(\text{g}) = \text{NH}_4\text{Cl}(\text{s})$ (1)
- 6) NH_3 , 2NH_3 2HCl , 8
 $(8 \cdot 22,4) / 44,8 = 4$ $= 405,2$ (1)

4. 7

- 1) $M = 29 / 0,96552 = 28$ / N_2 , CO , C_2H_4 . (1)
- 2) CO C_2H_4 :
 $2\text{CO} + \text{O}_2 = 2\text{CO}_2$; (1)

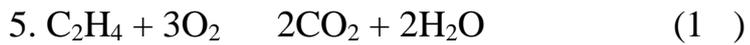
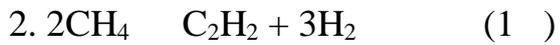


3)

2000°): (1)



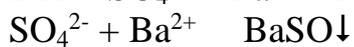
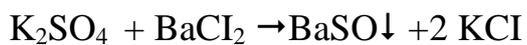
5. 7



6. 7

:

					.
		1	2	3	
I					2 - ;
II	BaCl ₂		-	-	1 -
		BaSO ₄			



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1)

2)

3)

4)

5)

- 1 _____.

- 1 _____.

, - 1 _____.

- 1 _____.

- 1 _____.