

1:1, ,

$n(gO):n(Fe_2O_3) = 4:1$ (1)

$n(Mg(NO_3)_2):n(Fe(NO_3)_3) = 2:1,$

$m(Mg(NO_3)_2):m(Fe(NO_3)_3) = 1,223:1$ (1)

: $(Mg(NO_3)_2) = 1,223\%$, $(Fe(NO_3)_3) = \%$ (1)

$(Mg(NO_3)_2) = \%$, $(Fe(NO_3)_3) = 0,818\%$ (1)

5

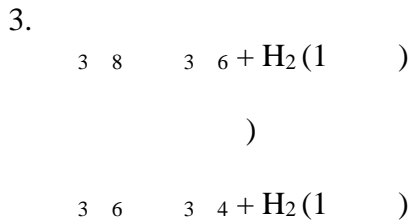
2. $(H_2 \quad n) = 2 + x + 16n /$
 $(K_2 \quad n) = 78 + x + 16n /$

:

$1() = \frac{1n}{2+x+16n} = 0,5854,$ $2() = \frac{1n}{7+x+16n} = 0,3038,$
 $= 32; n = 3$ (2)

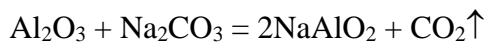
$SO_3(1)$, $2(O_3)(2)$ (1)

6



$44/2,2 = 20 /$,

5



$n(Al_2O_3) = 10,2/102 = 0,1$

$n(Na_2CO_3) = 21,2/106 = 0,2$ (, : $0,2 - 0,1 = 0,1$)

$n(NaAlO_2) = 0,2$

$m(CO_2) = 0,1 \cdot 44 = 4,4$ (1)

2)



$n(HCl) = 2n(Na_2CO_3) = 2 \cdot 0,1 = 0,2$ (1)

3)



n(HCl) = 4n(NaAlO₂) = 0,2 · 4 = 0,8 (1)

4)

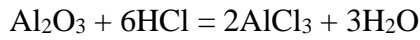
n(HCl) = 0,2 + 0,8 = 1

m(HCl) = 1 · 36,5 = 36,5

m(HCl) = 250 · 1,1 · 0,2 - 36,5 = 18,5

m(CO₂) = 0,1 · 44 = 4,4

w(HCl) = $\frac{m(H)}{m(p-pa)} = \frac{1,5}{1,2 + 2,2 + 2 \cdot 1,1 - 4,4 - 4,4} = 0,062, 6,2\% (1)$



HCl,

(2)

» (1)

7

5.

-) CuSO₄
-) Al(OH)₃
-) NH₄Cl
-) CO₂
-) K₂Cr₂O₇
-) HNO₃

) -) 1

- 1

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