

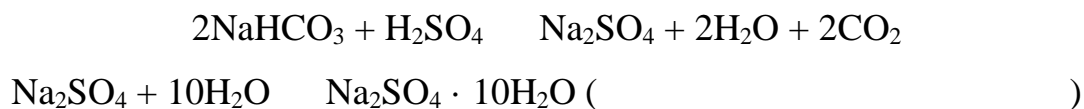
9 (2014)
9

1

1.1. :

	,	
$\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$. .	
KClO_3	.	
$\text{K}_3[\text{Fe}(\text{CN})_6]$.	(III)
H_2SO_5	.	
NaHCO_3		

1.2. :



2

1 , 500

HCl

$$n(\text{HCl}) = \frac{V(\text{HCl})}{V_m} = \frac{500}{22.4 /} = 22.3$$

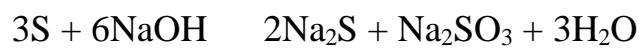
$$m(\text{HCl}) = n(\text{HCl}) \times M(\text{HCl}) = 22.3 \times 36.5 / = 815$$

$$m(H_2O) = \rho(H_2O) \times V(H_2O) = 1000 \text{ kg/m}^3 \times 1 \text{ m}^3 = 1000 \text{ kg}$$

$$m_{\text{total}} = m(H_2O) + m(HCl) = 1000 \text{ kg} + 815 \text{ kg} = 1815 \text{ kg}$$

$$\xi = \frac{m(HCl)}{m_{\text{total}}} \times 100\% = \frac{815}{1815} \times 100\% = 45\%$$

3



$$n(\text{H}_2) = V(\text{H}_2)/V_m = 3.36 / 22.4 = 0.15$$

0.1

, 2.7
 , 17.7

$$17.7 - 2.7 = 15.0 \quad x -$$

, 1.5x -

$$- 32 \cdot 1.5x = 48x. \quad 27x, \quad 27x + 48x = 15,$$

$$x = 0.2$$

$$0.2 \cdot 27 + 2.7 = 8.1$$

$$0.2 \cdot 48 = 9.6$$

$$8.1 / 17.7 = 0.46, \quad 46 \%$$

4

4.1. A - , - , - (II).

4.2. :



4.3. CuO Cu

, ,





1.	
(8)	16
	4
	20
2.	
	5
	5
	5
	5
	5
	25
3.	
Al ₂ S ₃	3
NaOH (3)	15
	4
	4
	4
	4
	30
4.	
, (3)	9
, (3)	12
CuO Cu	4
	25
	100