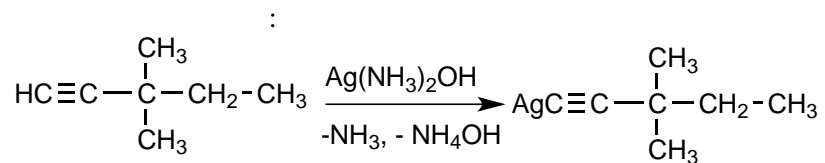


3,3-
3,3-

-1,4-
-1-

()

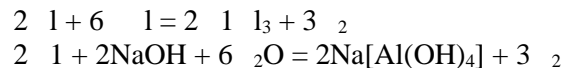


(8); - 2 (4); - 2 . - 25 .

- 4 (8); - 5 ,

10.1 (25)

: 100

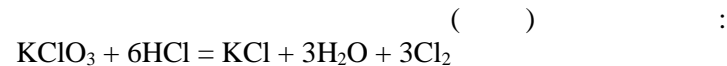


5%- 10 0,5 (10·0,05)

$$\begin{aligned} (1) &= 27 / \\ (1) &= 36,5 / \\ M(\text{NaOH}) &= 40 / \end{aligned}$$

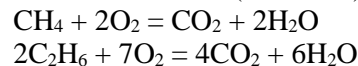
$$\begin{aligned} m(\text{HCl}) &= (0,1/27) \cdot 3 \cdot 36,5 = 0,406 \\ m(\text{NaOH}) &= (0,1/27) \cdot 40 = 0,148 \\ &1 \text{ NaOH} \end{aligned}$$

10.2 (30)



2 (8). - 30 . 4 (16), - 6 ,

10.3 (20)



() x y.

$$\begin{aligned} x + y &= 1/22,4 \\ 16x &= 30y \end{aligned}$$

$$x = 0,0291 ; y = 0,0155$$

$$\begin{aligned} n(\text{O}_2) &= 2 \cdot 0,0291 + 3,5 \cdot 0,0155 = 0,112 \\ V(\text{O}_2) &= 22,4 \cdot 0,112 = 2,52 \end{aligned}$$

() .

$$\begin{aligned} n(\text{CO}_2) &= 0,0291 + 2 \cdot 0,0155 = 0,0601 \\ n(\text{O}_2) &= 5/22,4 - 0,112 = 0,111 \\ m(\text{CO}_2) &= 0,0601 \cdot 44 = 2,64 \\ m(\text{O}_2) &= 0,111 \cdot 32 = 3,55 \end{aligned}$$

$$\begin{aligned} \text{CO}_2 &35,1\%, & 42,6\% \\ \text{O}_2 &64,9\%, & 57,4\% \end{aligned}$$

2 (4); - 4 (16). - 4 . - 20 .

10.4 (25)