

...

: 100

$$4. N() = 6,02 \cdot 10^{23} \cdot 0,002 = 1,2 \cdot 10^{21}$$

5 . - 20 .

8.1 (20)

, :

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20 .

8.2 (30)

- 1) $2\text{Li} + 2 \text{H}_2\text{O} = 2\text{LiOH} + \text{H}_2$
- 2) $2 \text{H}_2 + 5\text{O}_2 = 4 \text{H}_2\text{O} + 2 \text{H}_2\text{O}$
- 3) $\text{Fe}_2\text{O}_3 + 6 \text{HCl} = 2\text{FeCl}_3 + 3\text{H}_2\text{O}$
- 4) $\text{CuO} + \text{H}_2 = \text{Cu} + \text{H}_2\text{O}$
- 5) $\text{H}_2 + \text{O}_2 = \text{H}_2\text{O}$

6 . - 30 .

8.3 (30)

- , - ,

:

$$m(\text{CO}_2) = 44 ; m(\text{N}_2) = 28$$

$$+ 44 = 48,8$$

$$+ 28 = 45,6$$

$$= 40; = 0,2$$

$$m() = 48 - 40 = 8$$

$$() = 8 / 0,2 = 40 /$$

30 .

8.4 (20)

1. $m() = 1 \cdot 0,98 = 0,98$
2. $m() = 0,98 \cdot 0,1 = 0,098$
3. $n() = 0,098 : 46 = 0,002$