

9

9-1

$$M(\text{NaOH}) \qquad 60 \cdot 0,2 = 12$$

$$96 \cdot 0,125 = 12$$

$$m_{p-pa} = 96 - 60 = 36$$

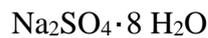
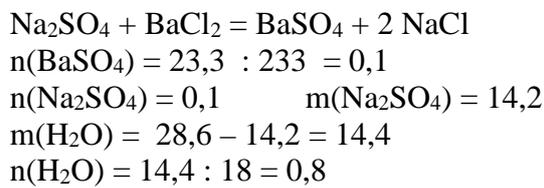
$$m(\text{NaOH}) = 36$$

$$m(\text{O}) = 36 - 4 = 32 \quad n(\text{O}) = 2$$

1. - 2
2. - 1
3. - 3

: 6

9-2



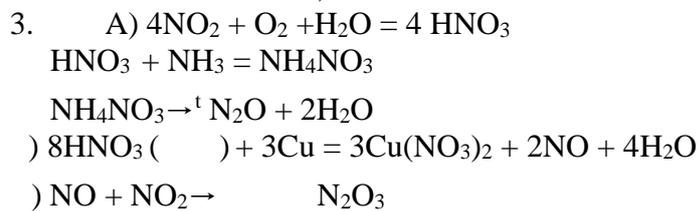
1. - 1
2. - 1
3. - 1

: 3

9-3



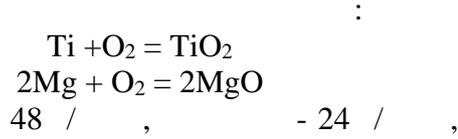
2.



1. - 1
2. - 2
3. - 2
4. N_2O - 2
5. 2 - 4

: 11

9-4



2, 1,5 48 () 22400 2.

$$x = \frac{1,5 \cdot 22400}{48} = 700$$

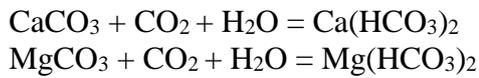
1. 1 - 2
2. - 3

: 5

9-5

$$\begin{aligned} C(\text{Ca}^{2+}) &= \frac{160}{1 \cdot 40} = 4 / \\ C(\text{Mg}^{2+}) &= \frac{48}{1 \cdot 24} = 2 / \\ C(\text{Na}^+) &= \frac{9.2}{1 \cdot 23} = 0.4 / \\ C(\text{HCO}_3^-) &= \frac{610}{1 \cdot 61} = 10 / \\ C(\text{SO}_4^{2-}) &= \frac{115.2}{1 \cdot 96} = 1.2 / \end{aligned}$$

Na₂SO₄; 1 MgSO₄; 1 MgCO₃; 2 CaCO₃ : 0,2



$$\begin{aligned} m(\text{Na}_2\text{SO}_4) &= 0,2 \cdot 142 / = 28,4 \\ m(\text{MgSO}_4) &= 1 \cdot 120 / = 120 \\ m(\text{MgCO}_3) &= 1 \cdot 84 / = 84 \\ m(\text{CaCO}_3) &= 2 \cdot 100 / = 200 \end{aligned}$$

1. 1 - 4
2. 2 - 8
3. - 3

: 15